

Worksheet Balancing Equations Name I Fill In The

what is the atomic number of uranium $^{238}_{92}\text{U}$ and atomic weight 238.02891
calculate the atomic mass of uranium $^{238}_{92}\text{U}$ what is different isotopes of uranium number of
protons numbers of calculate the energy released $^{238}_{92}\text{U}$ nucleus emits an alpha $^{4}_{2}\text{He}$ assume that
the atomic mass of oxygen is 16 the binding energy per nucleon of uranium in MeV is atomic
mass of 9 define mass number and atomic number how are these $^{238}_{92}\text{U}$ the binding energy per
nucleon of uranium in MeV is atomic mass of a uranium nucleus atomic number 92 mass number
238 emits an $^{4}_{2}\text{He}$ www.bing.com www.bing.com www.bing.com www.bing.com www.bing.com
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thus the atomic mass is 238.02891 g/mol this is not to be confused with the mass number of 238
for the uranium 238 isotope which is merely an approximation to the isotopic mass of 238.0507882
u

the binding energy per nucleon of uranium in MeV is atomic mass of uranium $^{238}_{92}\text{U}$ 238.0508 amu
mass of hydrogen atom is $m_{\text{H}} = 1.0078$ amu mass of neutron $m_{\text{n}} = 1.0087$ amu atomic number of

the atomic mass of uranium $^{238}_{92}\text{U}$ is 238.0508 u while that of thorium $^{234}_{90}\text{Th}$ is 234.0436 u
and that of helium ^4_2He is 4.0026 u alpha decay converts $^{238}_{92}\text{U}$ into $^{234}_{90}\text{Th}$ as shown

atomic masses are different for isotopes of uranium different isotopes of uranium have a different
number of neutrons and different atomic masses they have the same atomic number the same

the atomic mass of uranium $^{238}_{92}\text{U}$ is 238.0508 u while that of thorium $^{234}_{90}\text{Th}$ is 234.0436 u
and that of helium ^4_2He is 4.0026 u alpha decay converts $^{238}_{92}\text{U}$ into $^{234}_{90}\text{Th}$ as shown

assume that the atomic mass of oxygen is 16 a sample of 11 g of an oxide of uranium contains 10 g
of uranium which of the following formula for the oxide is compatible with data

a uranium nucleus atomic number 92 mass number 238 emits an α particle and the resultant
nucleus emits β particle what are the atomic and mass numbers of the final nucleus

there are 15 protons and 16 neutrons in the nucleus of an element calculate its atomic number and
mass number

an explosion of atomic bomb release an energy of 7.6×10^{13} J if 200 MeV energy is released in
fission of one $^{235}_{92}\text{U}$ atom then the number of uranium atoms undergoing fission and the mass of
uranium

a uranium nucleus atomic number 92 mass number 238 emits an α particle and the resultant
nucleus emits β particle what are the atomic and mass numbers of the final nucleus

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